

Precipitation using PolyPocket sleeves

(Belgium)

Background

In a precipitation reaction crystals of an insoluble salt are produced in solution. This is also referred to as settling out of solution. The precipitate is indicated by the symbol \downarrow or (s).

Safety

Photocopy this page (or similar) and place it in a polypocket for use during the experiment.

You will need...

- ✓ 1.2 M Na_2SO_4 solution
- ✓ 2 M CaCl_2 solution, a magnifying glass (optional)

Follow these steps

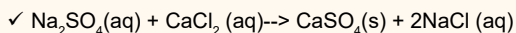
1. On the left circle, place one drop of 1.2 M Na_2SO_4 solution.
2. On the right circle, place a drop of 2 M CaCl_2 solution



3. Then fold paper to slide the two drops towards the large circle in the centre.
4. Observe

So what happened?

When ionic compounds are dissolved in water, they break apart into their constituent ions. In this case the ions re-arrange to form calcium sulfate (insoluble in water) and sodium chloride (soluble in water).



What next?

Explore the nature of covalent or ionic compounds, relative molecular mass and stoichiometry

